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2. Which, If Any, Of The Following Aqueous Mixtures

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Molecule,  $H_2O$  • Each O-H Bond Is Highly Polar,

Because Of The High Electronegativity Of The Oxygen

(N, O, F, And Cl Have High Values) • bond Angle Of

Water =  $105^\circ$  • due To The Bent Shape, The Jan 15th,

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Thus When We Consider HCl A Strong Acid In Aqueous

Medium. • Note, Because Of The Common Ion Effect, If

HCl Is Dissolved In Acetic Acid, The Extent Of Ioni Feb

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Simple Tri-atomic Molecule,  $H_2O$ . • Each O-H Bond Is

Highly Polar, Because Of The High Electronegativity Of

The Oxygen (N, O, F, And Cl Have High Values) • bond

Angle Of Water =  $105^\circ$ . Chapter 15 Water And

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