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33. $\text{KClO}_4 \rightarrow 2\text{O}_2 + \text{KCl}$ Actual yield = 1.45g

Theory yield = $5.95 \text{g KClO}_4 \times 1 \text{mol} / 122.55 \text{g} \times 2 \text{O}_2 \times 1 \text{KClO}_4 \times 32.00 \text{g} / 1 \text{mol} = 3.11 \text{g O}_2$
% Yield = $1.45 \text{g} / 3.11 \text{g} \times 100\% = 46.6\%$

34. $4\text{FeCO}_3 + \text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3 + 4\text{CO}$

2 Percentpurity=62.8%?gpureFeCO 3 20th, 2024.
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Chemistry Workbook Answers , Nakamichi 582 User
Manual , Walmart Application Answers , Acls
Certification Study Guide , Deltek Electronic Expense
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Manual , American English File 3 11th,
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Conversion Factor, We Determine The Mass That This

Number Of Moles Of SO₃ Has. The First Step

Resembles The Exercises We Did In Section 6.4 "Mole-

Mole Relationships In Chemical Reactions". As Usual,

We Start With The Quantity We Were Given: 3.59 Mol

Fe₂O₃ × 3 Mol SO₃ / 1 Mol Fe₂O₃ = 10.77 Mol SO₃ The

Mol Fe₂O₃ Units Cancel, Leaving Mol SO₃ ... 14th,

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