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1. STOICHIOMETRY INVOLVING ONLY PURE SUBSTANCES

(c) How Many Moles Of H2O Are Needed To Make 12 Moles Of NH3? (Ans. 36 Moles) (d) Calculate The Mass Of Ca(OH)2 That Can Be Formed From The Reaction Of 10.0 G Of Ca3N2 With An Excess Of H2O. Solution: Moles Ca3N2 = $10.0 \, \text{G}$ Ca3N2 X = $0.06743 \, \text{Mole}$ 1 Mole Ca3N2 148.3 G Ca3N2 Moles Ca(OH)2 = $0.06743 \, \text{Mole}$ Ca3N2 X = Feb 16th, 2024

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