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WORKSHEET- STOICHIOMETRY AND CHEMICAL FORMULA CALCUATIONS SET A: (Time Required, 1 Hour) 1. A Compound With The Formula, $BxH20O3$, Contains 36.14 % By Mass Oxygen. What Is The Value Of The Integer, X ? 1) Ans: $X = 6$ Feb 4th, 2024

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Chemistry 11 Stoichiometry Review Worksheet

Name: ___KEY___ 33. $\text{KClO}_4 \rightarrow 2\text{O}_2 + \text{KCl}$

Actual yield = 1.45g Theoretical yield = 5.95g $\text{KClO}_4 \times 1\text{mol}$

$122.55\text{g} \times 2\text{O}_2 \times 32.00\text{g} \times 1\text{mol} = 3.11\text{g O}_2$

% Yield = $\frac{1.45\text{g}}{3.11\text{g}} \times 100\% = 46.6\%$ 34. $4\text{FeCO}_3 + \text{O}_2$

$\rightarrow 2\text{Fe}_2\text{O}_3 + 4\text{CO}_2$

Percent purity = 62.8%? pure FeCO_3 Jan 5th, 2024

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