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Molar Mass =  $2 \times 12.011 + 6 \times 1.008 + 1 \times 15.999 = 46.069 \text{ G/mol}$   
 $\% \text{ C} = \frac{2 \times 12.011}{46.069} \times 100 = 52.14 \%$   
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1) How Many Moles Of Sodium Atoms Correspond To  
 $1.56 \times 10^{21}$  Atoms Of Sodium?  $1.56 \times 10^{21}$  Atoms Na X

1 Mol Na =  $2.59 \times 10^3$  Mol Na  $236.022 \times 10$  Atoms Na  
2) Determine The Mass In Grams Of Each Of The  
Following: A. 1.35 Mol Of Fe  $1.35 \text{ Mol Fe} \times 55.845 \text{ G Fe}$   
= 75.4 G Fe 1 Mol Fe B. 24.5 Mol O Jan 10th, 2024

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Chapter 3 In Brady, Russell, And Holum 's Chemistry,  
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Used In Sewage Treatment Plants. ° A. Construct A Pair  
Of Conversion Factors That Relate Moles Of Aluminum  
To Moles Of Sulfur For This Compound Feb 12th, 2024

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4 • ? H<sub>2</sub>O Weighing 24.95 g Is Heated Until All The Water Is Removed. The Anhydrate Weighs 15.95 g. What Is The Value Of ? In The Hydrated Salt? (MM CuSO<sub>4</sub>, 159.5 g/mol) A. 2 B. 3 C. 4 D. 5 27. The Simplest Formula For A Hydrocarbon That Is 20.0 Percent Hydrogen By M Jan 4th, 2024

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Percent Composition And Empirical Formulas 8) What Is The Mass Percent Of Each Element In The Following Compounds? A. CaCl<sub>2</sub> B. Fe<sub>2</sub>O<sub>3</sub> C. C<sub>6</sub>H<sub>10</sub>S<sub>2</sub>O<sub>9</sub> Calculate The Empirical Formulas Of Compounds Containing The Following Percentages Of Elements. Use The Molar Mass To Calc Mar 5th, 2024

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LESSON 1: Find The Gram Formula Mass Of The Following: (Show All Work) 1. MgO 5. Ca(OH)<sub>2</sub> 2. NaHCO<sub>3</sub> 6. CH<sub>4</sub> 3. C<sub>6</sub>H<sub>12</sub>O<sub>6</sub> 7. NH<sub>3</sub> 4. Al Feb 9th, 2024

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— Using The Molar Mass Of Each Substance, Convert The Following Quantities. A. 10.0 Mol Cr 520 G F. Mole On Grams, Grams On Moles Conversion Worksheet Units - 4 Moles And Stoichiometry Site 1 Of 2 Moles On Grams, Grams On Moles Conversion Worksheet What Are The Molecularthesis Of The Following Compounds? 1) NaOH 2) H 3PO 4 Apr 11th, 2024

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