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 Chemistry '14 Mr ...-2) Is A Polyatomic Ion, And The Sum Of All Oxidation Numbers
 In A Polyatomic Ion Equals The Charge Of The Polyatomic Ion. The Charge For (SO
 4-2) Is -2, So ALL Elements In (SO 4-2) Will Add To A -2 Oxidation Number. C.
 Oxygen As Stated Will Have An Oxidation Number Of -2, So The Four Oxygen Atoms
 Will Have A Total Of -8 Oxidation Number. D. As ... 1th, 2024.

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 Represents The Oxidation Reaction That Occurs Is A) $\text{HCl} + \text{KOH} \rightarrow \text{KCl} + \text{H}_2\text{O}$ B) 4
 $\text{HCl} + \text{MnO}_2 \rightarrow \text{MnCl}_2 + 2 \text{H}_2\text{O} + \text{Cl}_2$ C) $2 \text{HCl} + \text{CaCO}_3 \rightarrow \text{CaCl}_2 + \text{H}_2\text{O} + \text{CO}_2$ D) 2
 $\text{HCl} + \text{FeS} \rightarrow \text{FeCl}_2 + \text{H}_2\text{S}$ 21. Which Equation Represents An Oxidation-reduction
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Hydrogen-oxygen Fuel Cell, Energy Is Produced By The Overall Reaction Represented Above. (a) When The Fuel Cell Operates At 25°C And 1.00 atm For 78.0 Minutes, 0.0746 mol Of O_2 Is Consumed. Calculate The Amount Of H_2 Consumed. 1th, 2024.

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