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The NH_4Br is a salt that is soluble in water and so will dissociate into the following ions 100%. $\text{R NH}_4\text{Br (s)} \rightarrow \text{NH}_4^+$ (aq) + Br^- (aq) | 0.005 0 0 C -0.005 +0.005 +0.005 End 0 0.005 0.005 We are then left with NH_4^+ ions in solution. NH_4^+ is a weak acid and thus will reach an equilibrium state. Associate 5th, 2024

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