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Potassium Chlorate (KClO_3) Is Heated, It Decomposes To Form Gases KCl And Oxygen (O_2).
 $2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$
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33. $\text{KClO}_4 \rightarrow 2 \text{O}_2 + \text{KCl}$ Actual yield = 1.45g Theoretical yield = 5.95g
 $\text{KClO}_4 \times 1 \text{mol} = 122.55 \text{g} \times 2 = 245.1 \text{g}$
 $\text{KClO}_4 \times 32.00 \text{g} = 1 \text{mol} = 3.11 \text{g}$
 O_2 %Yield = $\frac{1.45 \text{g}}{3.11 \text{g}} \times 100\% = 46.6\%$
34. $4 \text{FeCO}_3 + \text{O}_2 \rightarrow 2 \text{Fe}_2\text{O}_3 + 4 \text{CO}_2$ Percent purity = 62.8%
?g pure FeCO_3 Mar 1th, 2024
Chemistry 12, Stoichiometry Review Problems, R. Janssen ...
In A Camping Stove? 2. How Many Moles Of O_2 Molecules Should Be Supplied To Burn 1 Mol Of CH_4 ...
Was At One Time Considered For Use As A Rocket Fuel. ... How Many Formula Units Of Calcium Oxide (CaO)
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