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Name:      KEY      33.  $\text{KClO}_4 \rightarrow 2\text{O}_2 + \text{KCl}$

Actual yield = 1.45g Theoretical yield = 5.95g  $\text{KClO}_4 \times 1\text{mol} = 122.55\text{g} \times 2 = 245.10\text{g}$   $\text{KClO}_4 \times 32.00\text{g} = 32.00\text{g} \times 1\text{mol} = 32.00\text{g}$

$\% \text{Yield} = \frac{1.45 \text{ g}}{3.11 \text{ g}} \times 100\% = 46.6\%$   
 $3\text{FeCO}_3 + \text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3 + 4\text{CO}_2$   
Percentpurity=62.8%?gpureFeCO<sub>3</sub> 22th, 2024

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11.4.1 Is Shown In Figure 12.2.1. We Can Use The  
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