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18 - 1 Chapter 18: Electrochemistry Oxidation States An Oxidation-reduction Reaction, Or Redox Reaction, Is One In Which Electrons Are Transferred. $2\text{Na} + \text{Cl}_2 \rightarrow 2\text{NaCl}$ Each Sodium Atom Is Losing One Electron To Form Na^+ $\text{Na} \rightarrow \text{Na}^+ + 1\text{e}^-$ This Loss Of Electrons Is Called Oxidation. Each Chlorine Atom Is Gaining 1 Electron

To Form Cl^- - $\text{Cl}_2 + 2\text{e}^-$ 1th, 2024

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Dr. Mattson, General Chemistry, Chm 205, Guide To Chapter 18. Electrochemistry 5
Read Section 18.8 Standard Cell Potentials And Equilibrium Constants. Learning
Objective 9: Use The Nernst Equation To Calculate The Equilibrium Constant, K. Do
Problems 13 And 14 At The End Of This Section. Do The Following End-of-chapter
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