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$= k[D]_1^x[B]_1^y$ $k[D]_2^x[B]_2^y$ Equation 5 In This Equation k Cancels Out. $[D]_1 = [D]_2$ Because We Are Starting Off With The Same Initial Concentration Of A In Each Trial. The Order Of Species D Which Is x Is Also Constant. Canceling Terms We Have Left: $Rate_1 = Rate_2 = k[B]_1^y$ $k[B]_2^y$ Equation 6 Consider The Following Chemical Reaction: Apr 7th, 2024

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Chemical Identity - What Exactly Is Reacting? Usually Ions Of Opposite Charge React Very Rapidly. Also, The More Bonds Between Reacting Atoms In A Molecule, The Slower The Reaction Rate. Why? More Energy Is Required To Separate The Molecule Into Its "bits". Substances With Strong Bonds (larger Bond Energies) Will React Much More Slowly. May 12th, 2024

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So, To Account For Effective Collisions, Another Factor, P Called The Probability Or Steric Factor Is Introduced. So, Rate = $PZ \cdot AB \cdot e^{-E_a/RT}$ The Activated Complex Theory Or Transition State Theory Reactants \rightleftharpoons Activated Complex \rightarrow Products This Theory Is Based On The Fact That May 14th, 2024

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