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87 The Total Number Of Moles,  $N$ , Is The Sum Of The Number Of Moles Of All The Components Of The Mixture:  $N = N_1 + N_2 + \dots$  It Follows That The Sum Of All The Mole Fractions Of The Mixture Must Also Equal 1.  $X_1 + X_2 + \dots = 1$  The Mass Of Component  $i$  In A Mixture Is The Product Of The Number Of Moles Of  $i$  And Its Molecular Weight,  $M_i$ . The Mass Of The Mixture Is Therefore The Sum,  $M = N_1M_1 + N_2M_2 + \dots$  Mar 11th, 2024

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Table C.2 (K) Nitrogen ( $N_2$ ) 100 150 200 250 300 350 450 550 700 1,000 1,100 1,200 1,300 Oxygen ( Apr 6th, 2024

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