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$$(\text{mol/L}) = 12.2 \frac{P(\text{atm})}{T-1(\text{K})}$$
 Where P And T Are The Absolute Pressure And Temperature Of The Gas. The Following Data Are Taken For The Feed Gas: T=750F, P=150 Psig, H=210 Mm, R=82.4 And For The Outlet Gas, R=11.6. Determine The (a) Mass Fraction Of SO Jan 3th, 2024
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