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Greater Than 1, And A Volume V_i . Both Inlet And Outlet Ports Are Turned On At Time $T=0$. The Inlet Port Flows With A Volumetric Flow Rate, V_0 , And A Concentration, C_{A0} . The Outlet Port Flows With A Volumetric Flowrate V_0/a , Where a Is The Sum Of The Stoichiometric Coefficients (See Table 10.1), That Integral Can Be Evaluated Using A Partial Fraction Decomposition Of The Integrand, Ultimately Allowing One To Solve For B As A Function Of T . The Solution Of Eq. 5 Is $B = \frac{C_{A0} V_0}{k} \left(1 - \frac{1}{1 + kT} \right)$; (6) Where C_{A0} Is A Constant Whose Value Depends On The Initial Conditions. Knowing The Reaction Order Can Help Chemists To Predict How The Reaction Speed Will Change Over Time, As Well Help Them To Know How Much Or Little To Change The Experimental Conditions That Influence The Rate Of The Reaction. The Iodine Clock Reaction Is A Classical Experiment Clock Demonstration. May 10th, 2024.

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