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$N\text{O} + n\text{e}^- \rightleftharpoons \text{NRR}$ (2.1.36) The Free Energy Of The Cell Reaction Is Given By, $\Delta G =$
 $\Delta G^0 + RT \ln \frac{a_{\text{R}}}{a_{\text{O}}}$ (2.1.38) Which Since $\Delta G = -nF E_{\text{rxn}}$
And $a_{\text{H}^+} = 1$ Becomes, $E = E^0 + \frac{RT}{nF} \ln \frac{a_{\text{R}}}{a_{\text{O}}}$. (2.1.40) This
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