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$N_2O + ne \rightarrow N_2$  (2.1.36) The Free Energy Of The Cell Reaction Is Given By,  $\Delta G = \Delta G^0 + RT \ln \frac{a_{R_1} a_{R_2} \dots a_{R_n}}{a_{O_1} a_{O_2} \dots a_{O_n}}$  (2.1.38) Which Since  $\Delta G = -nFE_{rxn}$  And  $\Delta G^0 = -nFE^0$  Becomes,  $E = E^0 + \frac{RT}{nF} \ln \frac{a_{R_1} a_{R_2} \dots a_{R_n}}{a_{O_1} a_{O_2} \dots a_{O_n}}$  (2.1.40) This Expression Is Called The Nernst Equation. This Relation Can Feb 20th, 2024

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