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FUELS AND COMBUSTION 3.1 Introduction To Combustion⁸⁷ The Total Number Of Moles, N , Is The Sum Of The Number Of Moles Of All The Components Of The Mixture: $N = N_1 + N_2 + \dots$ It Follows That The Sum Of All The Mole Fractions Of The Mixture Must Also Equal 1. $X_1 + X_2 + \dots = 1$ The Mass Of Component i In A Mixture Is The Product Of The Number Of Moles Of i And Its Molecular Weight, M_i . The Mass Of The Mixture Is Therefore The Sum, $M = N_1M_1 + N_2M_2 + \dots$

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