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Show Your Work, Including Proper Units, To Earn Full Credit.  
1. Silver And Nitric Acid React According To The

Following Balanced Equation:  $3 \text{Ag}(s) + 4 \text{HNO}_3(aq) \rightarrow 3 \text{AgNO}_3(aq) + 2 \text{H}_2\text{O}(l) + \text{NO} \dots$  May 3th, 2024  
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1. Silver And Nitric Acid React According To The Following Balanced Equation:  
 $3 \text{Ag} + 4 \text{HNO}_3 \rightarrow 3 \text{AgNO}_3 + \text{NO} + 2 \text{H}_2\text{O}$   
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KEY 9) 2 24 2 2 23 2 2 2 4.63 X 10 molecules | 1 Mol | 6.02 X 10 Molecules | 1 Mol Cl 1mol 71 G Cl Cl X 546 G Cl 10) 292 G Ag 1 Mol Ag 108 G Ag 1 Mol Cu 1 Mol Ag 63.5 G Cu  
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Directions: Solve Each Of The Following Problems. Show Your Work, Including Proper Units  
1.  $\_\_ \text{CaCl}_2 + \_\_ \text{AgNO}_3 \rightarrow \_\_ \text{Ca}(\text{NO}_3)_2 + \_\_ \text{AgCl}$   
How Many Grams Of Silver Chloride Are Produced When 45 G Of Calcium Chloride React With Excess Silver ...  
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Name: \_\_\_KEY\_\_\_ 33.  $\text{KClO}_4 \rightarrow 2\text{O}_2 + \text{KCl}$

Actual yield = 1.45g Theoretical yield = 5.95g  $\text{KClO}_4 \times 1 \text{ mol}$

$122.55 \text{ g} \times 2 \text{ O}_2 \times 1 \text{ KClO}_4 \times 32.00 \text{ g} \times 1 \text{ mol} = 3.11 \text{ g O}_2$

% Yield =  $\frac{1.45 \text{ g}}{3.11 \text{ g}} \times 100\% = 46.6\%$  34.  $4\text{FeCO}_3 + \text{O}_2$

$\rightarrow 2\text{Fe}_2\text{O}_3 + 4\text{CO}_2$

Percent purity = 62.8%? pure  $\text{FeCO}_3$  May 6th, 2024

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