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 $\text{Actual yield} = 1.45\text{g}$ $\text{Theoyield} = 5.95\text{g}$
 $\text{KClO}_4 \times 1\text{mol} = 122.55\text{g}$
 $2\text{O}_2 \times 32.00\text{g} = 64.00\text{g}$
 $\text{O}_2 = 3.11\text{g}$
 $\% \text{Yield} = \frac{1.45\text{g}}{3.11\text{g}} \times 100\% = 46.6\%$
 34. $4\text{FeCO}_3 + \text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3 + 4\text{CO}_2$
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