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CHAPTER 6 THERMOCHEMISTRY

162 CHAPTER 6: THERMOCHEMISTRY To Convert The Answer To Joules, We Write: $101.3 \text{ J} \cdot 0.18 \text{ L} \cdot \text{Atm} = 18.2 \text{ J}$ $1 \text{ L} \cdot \text{Atm} = 101.3 \text{ J}$ $18.2 \text{ J} \cdot \frac{1 \text{ L} \cdot \text{Atm}}{101.3 \text{ J}} = 0.18 \text{ L} \cdot \text{Atm}$ -18 J 6.17 An Expansion Implies An Increase In Volume, Therefore W Must Be -325 J (see The Defining Equation For Pressure-volume Work.) If The System Absorbs Heat, Q Must Be $+127 \text{ J}$. The Change In Energy (internal 3th, 2024

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1 Chapter 3: Thermochemistry 1. Which Is A Unit Of Energy? A. Pascal B. Newton *

C. Joule D. Watt E. Ampere 2. For A Chemical Reaction, Where The Internal Energy Is ... 27th, 2024

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To The sum Of energy in Form Heat Exchanged Between System And Surroundings (q) And The Work Done On The System (w). $\Delta E_{\text{system}} = Q + W$ But The Laws Of Thermodynamics Are Far More Profound Than Simply Describing The Efficiency Of Steam Engines. T 12th, 2024

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