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Near The Oxygen End Of The Molecule And Partial
Positive Charge Near The Hydrogen Atoms. Thus, The
Negative End Of A Water Molecule Is Strongly

Attracted To ... 2th, 2024Chapter 3

StoichiometryChapter 3 Stoichiometry 3-7 EXAMPLE

PROBLEM: Determine The Molar Mass Of A Compound

Calculate The Molar Mass For Each Of The Following

Compounds: (a) 2-Propanol, $\text{CH}_3\text{CH}(\text{OH})\text{CH}_3$ (b) Iron(II)

Phosphate, $\text{Fe}_3(\text{PO}_4)_2$ SOLUTION: You Are Asked To

Calculate The Molar Mass Of A Compound. You Are

Given The Compound's Formula. 3th, 2024.

CHAPTER 2: Atoms, Molecules And Stoichiometry2.5

Stoichiometry And Equations Stoichiometry 1)

Stoichiometry Is The Proportion Of Things Either

Reacting Or Combining. 2) In Compounds, It Refers To

The Ratio In Which The Atoms Are Combined Together.

For Example, Water, H_2O Has A Stoichiometry Of 2

Hydrogen To 1 Oxygen. 3) It Also Refers To The

Reacting Proportions In A Chemical Equation. 2th,

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