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$\text{NH}_3 + \frac{7}{2}\text{O}_2 \rightarrow 2\text{NO}_2 + 3\text{H}_2\text{O}$. The Question Asks

About The Amount Of Oxygen Reacting With ONE Mole

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Agenda: 1) Go Over Quest 2) Introduce Stoichiometry-S'more Activity HW: S'More Worksheet #2
Tuesday, February 14
Agenda: Lab -Introduction To Stoichiometry HW: 1) Lab Calculations 2) Read P 237-240
Wednesday, February 15
Agenda: 1) Go Over Lab 2) Worksheet #3 HW: 1) Read P 242 2) Video Notes
Thursday, Feb 2th, 2024.
Stoichiometry Test Review March 21/225.
Given The Incomplete Equation: $2 \text{N}_2\text{O}_5(\text{g}) \rightarrow$ Which Set Of Products Completes And Balances The Incomplete Equation? 6.
Base Your Answer To The Following Question On Given The Balanced Equation Representing A Reaction: $\text{C}_3\text{H}_8(\text{g}) + 5\text{O}_2(\text{g}) \rightarrow 3\text{CO}_2(\text{g}) + 4\text{H}_2\text{O}(\text{g})$
What Is The Total Number Of Moles Of $\text{O}_2(\text{g})$
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